

IN THE SPECIFICATION:

The specification as amended below with replacement paragraphs shows added text with underlining and deleted text with ~~striethrough~~.

Please REPLACE the paragraph beginning at page 5, line 18 (paragraph [0023]), with the following paragraph:

[0023] The carbon compound may be present in the surface of the lithium transition metal complex oxide in the form of either an oxide or a carbonate compound, such as CO, CO₂, Li₂CO₃, and the like. The carbon compound is derived from source materials used in the preparation of the lithium transition metal complex oxide, externally supplied CO₂ or O₂, or a reaction product thereof. Although the carbon compound is described as being adsorbed into the surface of the lithium transition metal complex oxide, it may be linked to the surface of the lithium transition metal complex oxide by a chemical bond such as a covalent bond. Therefore, the description “adsorption of the carbon compound into the surface of the lithium transition metal complex oxide” throughout the specification must not be interpreted as excluding a link to the lithium transition metal complex oxide by a chemical bond.

Please REPLACE the paragraph beginning at page 8, line 22 (paragraph [0039]), with the following paragraph:

[0039] Meanwhile, any separator commonly used in the manufacture of lithium batteries may be used for a lithium battery according to an embodiment of the present invention. However, preferred materials for the separator allow electrolyte ions to migrate with less resistance and have the ability to retain a larger amount of electrolytic solution. Specific examples of such separator materials include a glass fiber, polyester, TEFLON (polytetrafluoroethylene), polyethylene, polypropylene, polytetrafluoroethylene (PTFE), and a combination of the foregoing materials, which may be in non-woven fabric or woven fabric form. In particular, a separator made of polyethylene, polypropylene, or the like, which can be rolled, may be used for a lithium ion battery, and a separator that may retain a larger amount of organic electrolytic solution may be used for a lithium ion polymer battery. These separators may be manufactured as follows.

Please REPLACE the paragraph beginning at page 10, line 2 (paragraph [0046]), with the following paragraph:

[0046] 96% by weight of LiCoO_2 , 2% by weight of polyvinylidene fluoride (PVDF) as a binder, and 2% by weight of a carbonaceous conducting agent that facilitates migration of electrons were mixed together, and 100 mL of N-methylpyrrolidone (NMP) and ceramic balls were added to the mixture and thoroughly mixed in a 200-mL plastic bottle for about 10 hours. A cathode was cast from the mixture on a 15 μm -thick ~~aluminium~~-aluminum foil using a 250 μm -spaced doctor blade, dried in an oven at 110 °C for about 12 hours to fully evaporate the NMP, roll pressed, and cut to a predetermined size to manufacture a cathode plate having a thickness of 95 μm .